phthalimid. A portion of the crystals heated with potassium hydroxide went into solution with evolution of ammonia. Another portion of the crystals were covered with concentrated ammonia and allowed to stand for some time. They were soon converted into microscopic crystals of phthalamid

$$(=C_6H_4(CONH_2)_2).$$

These crystals were filtered off, washed, and dried. They melted at 217.5° (uncorrected) with an evolution of ammonia, which began at about 200°. The phthalamid was further proved by its insolubility in cold water, alcohol, and ether, and by boiling it with water it was decomposed, giving off ammonia and on cooling phthalimid, melting at 230° C., crystallized out.

The results of these tests show conclusively that the product is phthalimid and that when it is made by the action of equal molecules of acid and nitril the yield is large. The reaction works comparatively readily, and at a much lower temperature than was needed to affect the reactions recorded by Colby and Dodge. It is highly probable that with slight changes of conditions any one of a variety of nitrils would give the same result. I hope to report further experiments with phthalic acid and other dibasic aromatic acids at a later day.

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DETERMINATION OF SULPHURIC ACID.

BY N. J. LANE.

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SOME months ago, before hearing of the controversy between Dr. Lunge and Mr. Gladding, some experiments were made on this subject, the results of which sustain Mr. Gladding's case. The determinations were made on nearly normal sulphuric acid to establish its strength with the following results:

	Barium chloride added suddenly.	Barium chloride added by drops.
1. Sulphuric acid	50.03	49.23
2. " "	49.90	49.32
3. " "	50.14	....

And the average of several practically identical titrations on C. P. sodium carbonate gave sulphuric acid 49.33.

The above results were obtained with the greatest care, and every precaution used to insure accuracy. This, in my opinion, conclusively proves the accuracy of Mr. Gladding's statements.